PASSAGGIO DA $\mu g/m^3$ a PPM di H_2S in aria -

$$\frac{1\ \mu g}{m^3} = \frac{(1\mu g \rightarrow 10^{-6} g)}{PM_{H_2S}\ in\ g} * \text{VOLUME MOLARE} = \frac{}{m^3} =$$

$$\frac{10^{-6} g}{PM_{H_2S}\ in\ g} = N^\circ \text{ molecole in } 1\mu g$$

$$\frac{10^{-6} g}{PM_{H_2S}\ g} * 22.414 lit = \text{Volume occupato dal quale moli espresse in litri}$$

$$\frac{10^{-6} g}{34 g} * 22.414\ lt = \frac{}{1000\ litri} = 0.659 \cdot 10^{-3} (PPM)$$

10^{-6}

$$\frac{1\ mg}{m^3} = \frac{0.659\ litri}{10^{+6} \cdot 10^{+3}\ litri} =$$

$$= 0{,}659 \cdot 10^{-3}\ litri \text{ su } 10^{6}\ litri$$

$$= 0{,}659 \cdot 10^{-3}\ ppm$$